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Mechanism of orthophosphate ($\text{PO}_4\text{-P}$) adsorption onto different biochars

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ABSTRACT

The adsorption mechanisms of phosphate ($\text{PO}_4\text{-P}$) onto cocoa pod husk (CP), corn cob (CC), rice husk (RH) and palm kernel shell (PK) biochar pyrolyzed at 300 °C and 650 °C were investigated. A series of batch experiments were undertaken to assess the effects of contact time and pH. Results show that $\text{PO}_4\text{-P}$ adsorption equilibria for the biochar types was within 6–15 h, being rapid in the 300 °C-biochar types. The equilibrium pH for maximum $\text{PO}_4\text{-P}$ adsorption varied among biochar types, ranging from of 2.6 to 4.8 and increasing with decreasing $\text{PO}_4\text{-P}$ adsorption. Pseudo-second-order and Elovich models explained the adsorption data well indicating a chemisorption process on heterogeneous biochar surface. $\text{PO}_4\text{-P}$ adsorption was controlled initially by intraparticle diffusion and subsequently by chemisorption. Per the properties of the biochars (FTIR and elemental composition) and pH (equilibrium pH and ΔpH), $\text{PO}_4\text{-P}$ was adsorbed through electrostatic attraction, surface precipitation and ligand exchange, and the relative importance of these processes differed among the biochar types. Biochar types (PK300, PK650, CP300, CP650, RH650 and CC650) that adsorbed $\text{PO}_4\text{-P}$ through surface precipitation and ligand exchange reactions can be used to remove $\text{PO}_4\text{-P}$ from wastewater since $\text{PO}_4\text{-P}$ is strongly adsorbed, controlling $\text{PO}_4\text{-P}$ enrichment of water bodies.

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1. Introduction

Adsorption of phosphate ($\text{PO}_4\text{-P}$) by organic material is of relevance in controlling the transport, mobility and enrichment of $\text{PO}_4\text{-P}$ in aquatic environment. Although $\text{PO}_4\text{-P}$ is an essential nutrient for the growth of plants and other living organisms, it can also serve as an environmental pollutant (Dodds et al., 2008; Almeelbi and Bezbaruah, 2012). About 0.02 mg L⁻¹ of dissolved $\text{PO}_4\text{-P}$ in water bodies is considered to have potential risk to the proliferation of algal growth (USEPA, 1995). The increased $\text{PO}_4\text{-P}$ in water bodies as a result of inappropriate and frequent discharge of waste water stimulates excessive growth of phytoplankton and algae which in turn decrease the quality of drinking water (Dodds et al., 2008). Thus, in order to reduce the negative effects of overloading water bodies with $\text{PO}_4\text{-P}$, it is necessary to assess various strategies and to evaluate the $\text{PO}_4\text{-P}$ removal effectiveness of adsorbents that could be exploited for use as waste water cleansing agents of $\text{PO}_4\text{-P}$ ions prior to wastewater discharge into natural water bodies (Biswas et al., 2008).

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Table 1

Properties of corn cob biochar, palm kernel shell biochar, cocoa pod husk biochar and rice husk biochar produced at 300 °C and 650 °C.

Biochar	Cocoa pod husk		Rice husk		Corn cob		Palm kernel shell	
	300 °C	650 °C	300 °C	650 °C	300 °C	650 °C	300 °C	650 °C
Abbreviation	CP300	CP650	RH300	RH650	CC300	CC650	PK300	PK650
pH (H ₂ O)	8.50	10.21	7.11	9.50	8.90	10.30	5.51	6.70
C (%)	57.94	61.96	41.58	42.55	72.98	78.07	81.99	89.58
CEC (cmol ₊ kg ⁻¹)	63.50	53.00	49.40	38.00	38.02	30.10	27.90	22.10
Total P (g kg ⁻¹)	2.45	3.16	1.19	1.52	1.54	1.94	0.17	0.25
Ca (g kg ⁻¹)	9.60	11.71	2.07	3.29	3.32	4.19	1.95	2.94
Mg (g kg ⁻¹)	7.23	8.86	0.91	1.27	2.43	3.43	0.27	0.33
K (g kg ⁻¹)	0.17	0.22	0.19	1.52	1.54	1.94	0.17	0.25
Na (g kg ⁻¹)	41.33	58.94	5.34	8.19	24.71	30.80	0.45	0.58
Si (g kg ⁻¹)	8.91	28.36	138.53	194.49	13.94	14.44	9.74	14.41
Total N (g kg ⁻¹)	3.91	2.73	1.77	1.21	2.93	1.91	3.82	2.12
Al (g kg ⁻¹)	0.46	0.99	1.26	1.88	0.57	1.10	1.04	2.11
Fe (g kg ⁻¹)	0.39	0.65	1.25	1.66	0.62	0.95	0.68	1.49
Carboxylic (mmol g ⁻¹)	0.55	0.09	0.38	0.10	0.47	0.06	0.21	0.05
Phenolic (mmol g ⁻¹)	1.22	0.46	1.01	0.23	0.71	0.32	1.32	0.65
Lactonic (mmol g ⁻¹)	0.27	0.03	0.19	0.00	0.11	0.04	0.08	0.00
Total acid group (mmol g ⁻¹)	2.12	0.94	2.04	0.60	2.10	0.97	1.82	0.80

To treat PO₄-P contaminated water, a number of technological approaches have been deployed such as electrodialysis, reverse osmosis, chemical precipitation, membrane technologies, biosorption and ion exchange (Demirbas, 2008). The less expensive, technologically and environmentally friendly among the aforementioned methods is biosorption technique, and biochar is a common biosorbent. Biochar is a carbon-rich, fine grained, and porous material. It is usually produced by thermal decomposition of biomass under limited oxygen conditions at temperatures <850 °C (Lehmann, 2007). It has gained a lot of attention in recent times not only for its role in soil fertility improvement and carbon sequestration but also as a low-cost adsorbent for wastewater treatment (Yao et al., 2012; Zhang et al., 2016). Reports have indicated that biochar has a strong ability to remove chemical contaminants in water including PO₄-P, organic contaminants and heavy metals (Kasozzi et al., 2010; Yao et al. 2013; Zhang et al., 2016). Due to its highly-porous structure, mineral contents (e.g. CaCO₃, Fe and Al) and functional groups (e.g. phenolic, carboxylic and amino groups), biochar has a high affinity for PO₄-P (Li et al., 2017). However, the adsorptive properties of biochar depends largely on the feedstock and pyrolysis temperature (Zhao et al., 2013).

Adsorption kinetics expresses the time-dependent removal of adsorbate by an adsorbent in an aqueous solution and its application is relevant for elucidating adsorption mechanism (Wahab et al., 2011; Jung et al., 2015; Zhang et al., 2015). Using chemical adsorption kinetic models (e.g. pseudo-first-order, pseudo-second-order, Elovich and Johnson-Mehl-Avrami), the mechanism of PO₄-P adsorption by biochar has been reported to be due to chemisorption (Yao et al., 2011; Wahab et al., 2011; Jung et al., 2015; Zhang et al., 2016). However, most of these studies did not focus on biochar at different pyrolysis temperatures. It is, therefore, important to clarify the magnitude of the chemical interactions between PO₄-P and biochar at different pyrolysis temperatures as a function of time and pH in a chemical system. In the present study, PO₄-P adsorption dynamics on four biochar types at two pyrolysis temperatures were investigated. The objectives of the study were to (1) determine contact time and pH effect on PO₄-P adsorption and (2) to elucidate possible mechanism for PO₄-P adsorption onto biochar.

2. Materials and methods

2.1. Biochar

The four biochar types produced at two different pyrolysis temperatures (300 °C and 650 °C) from cocoa pod husk (CP) (*Theobroma cacao*), corn cob (CC) (*Zea mays*), rice husk (RH) (*Oryza sativa*) and palm kernel shell (PK) (*Elaeis guineensis*) were obtained from the Soil Research Institute of the Council for Scientific and Industrial Research (CSIR), Ghana. All biochars were finely ground to <1 mm using a mortar and pestle and stored in air-tight bags until characterization and kinetics studies. The CP, CC, RH and PK feedstocks pyrolyzed at 300 °C and 650 °C are herein after designated as CP300, CP650, CC300, CC650, RH300, RH650, PK300 and PK650, respectively. Detailed description of the biochar types used for the study are shown in Table 1. Briefly, the pH, total P, total C, the concentration of the base cations (Ca, Mg, K and Na) and heavy metals (Fe and Al) increased with increasing pyrolysis temperature with the exception of the of total N and cation exchange capacity (CEC) which decreased otherwise.

2.2. Fourier transform infrared (FTIR) analysis and Boehm titration

The functional groups on the biochar types were examined using photoacoustic spectroscopy (PAS)-FTIR. In short, the spectra were recorded using a Nicolet 6700 (Thermo Scientific, USA) spectrometer equipped with a PA-301 photoacoustic

detector (Gasera Ltd., Finland). For each sample, 128 scans in the mid-infrared region between 4000 and 500 cm^{-1} at a resolution of 4 cm^{-1} were recorded and averaged. Boehm titration method was used to determine surface acid functional groups on the biochar types (Boehm, 1994; Usman et al., 2015). Shortly, after adjusting the pH of biochar (1.0 g) and deionized water (40 cm^3) suspension to 5.0, it was left undisturbed for 7 days, dried at 60 $^{\circ}\text{C}$, and homogenized. To each of 10 mL of 0.05 M NaOH, 0.10 g of the homogenized biochar samples was added. The mixtures, along with a control sample (i.e. solution without biochar), were shaken for 24 h. An aliquot of 5 mL of the filtrates was then pre-neutralized with 10 mL of 0.05 M HCl and titrated to pH 7.0 by adding 0.05 M NaOH solution at a rate of 0.5 mL min^{-1} controlled by a TIM automatic titration system (Radiometer, France). The calculation for surface acid functional groups was done on the basis that 0.05 M NaOH neutralizes all acidic groups including carboxyl, lactonic, and phenolic groups (Boehm, 1994; Jiang et al., 2015).

2.3. Kinetics analysis and pH effect on $\text{PO}_4\text{-P}$ adsorption

Kinetics study was carried out to determine the equilibrium time of $\text{PO}_4\text{-P}$ adsorption onto the biochar types. $\text{PO}_4\text{-P}$ solution was prepared by dissolving potassium dihydrogen phosphate (KH_2PO_4) in 10 mM potassium chloride (KCl). The $\text{PO}_4\text{-P}$ kinetics of each biochar were examined by mixing 0.5 g of biochar with 30 mL of 100 mg L^{-1} $\text{PO}_4\text{-P}$ solution in 50 mL centrifuge tubes. The tubes were shaken on an end-to-end shaker at 120 rpm for 24 h at room temperature (22 ± 0.5 $^{\circ}\text{C}$). A tube was withdrawn after the shaking was done for 1, 3, 5, 9, 15 and 24 h. The suspension was then centrifuged (3500 rpm, 15 min) and the resulting supernatant was filtered through a 0.45 μm filter paper. The filtrate was analyzed for $\text{PO}_4\text{-P}$ by the colorimetric molybdenum-blue method (John, 1970).

The effect of solution pH on $\text{PO}_4\text{-P}$ adsorption onto biochars was studied in an initial pH range of 2.5, 3.5, 4.5, 5.5, 6.5, 7.5 and 8.5. The pH was adjusted using 4 M HCl or NaOH. The ratio of biochar to the initial solution $\text{PO}_4\text{-P}$ concentration was the same as the kinetics experiment. Furthermore, for each pH, a blank (0.5 g biochar per 30 mL deionized water without $\text{PO}_4\text{-P}$ solution) was also included. The centrifuge tubes were shaken on an end-to-end shaker at 120 rpm for 24 h at room temperature (22 ± 0.5 $^{\circ}\text{C}$). The suspension was then centrifuged (3500 rpm, 15 min) and the resulting supernatant was filtered through a 0.45 μm filter paper. The filtrate was analyzed for $\text{PO}_4\text{-P}$ by the colorimetric molybdenum-blue method (John, 1970) and equilibrium (final) pH was also measured. The initial aqueous $\text{PO}_4\text{-P}$ concentration C_i (mg L^{-1}) and the equilibrium $\text{PO}_4\text{-P}$ concentration C_e (mg L^{-1}) were measured and the $\text{PO}_4\text{-P}$ adsorbed (q_t) was computed from the mass balance equation (Eq. (1)):

$$q_t = \frac{(C_i - C_e)V}{M} \quad (1)$$

Here, V is the volume of the aqueous solution (L) and M is the dry weight in grams of biochar (adsorbent).

The kinetics experiments and the effect of equilibrium pH on $\text{PO}_4\text{-P}$ adsorption experiments were repeated three times. The average data and standard errors are reported.

2.4. Statistical analysis

The coefficient of determination (R^2), standard error and other statistics, as well as fitting all kinetics data were done using Origin Pro 8.5 and RStudio v.3.4.2. The standard errors of the fitted coefficients were within 95% confidence intervals. Error bars were used to represent standard errors of triple determinations.

3. Results and discussion

3.1. Biochar FTIR and Boehm titration

FTIR spectrum of CC, PK, CP and RH at 300 $^{\circ}\text{C}$ and 650 $^{\circ}\text{C}$ revealed the information about adsorption process (Fig. 1). The bands at 1720/1396, 3400 and 2925 cm^{-1} prove the existence of carboxylic C ($-\text{COOH}$), phenolic O-H and aliphatic $-\text{CH}_2$ stretch, respectively (Chun et al., 2004; Bourke et al., 2007; Brewer et al., 2009). These vibrational bands were relatively intensified at 300 $^{\circ}\text{C}$ -biochar types as compared to their counterparts. The decrease in peaks 3400 cm^{-1} , 1396 cm^{-1} and 1720 cm^{-1} at 650 $^{\circ}\text{C}$ suggests the decrease in acid functional groups at high pyrolysis temperature (Brewer et al., 2009). Usman et al. (2015) and Chun et al. (2004) also reported a decrease in $-\text{COOH}$ groups (acid functional group) with increasing pyrolysis temperature and even disappearance at high temperature. A well-stretched band centered around 800 cm^{-1} and 1100 cm^{-1} correspond to silicates (Bourke et al., 2007). Another noticeable band was seen at 875 cm^{-1} , which is assigned carbonate (Brewer et al., 2009). Results from the Boehm titration were used to further quantify the acid functional groups on the biochar types at the two temperature regimes (300 $^{\circ}\text{C}$ and 650 $^{\circ}\text{C}$) (Table 1). Obviously, the surface acid functional groups (carboxylic, phenolic and lactonic groups) varied with feedstock and pyrolysis temperature. The dominance of the total surface acid functional groups on the biochar types were in the order of CC ($0.97\text{--}2.10$ mmol g^{-1}) > CP ($0.94\text{--}2.12$ mmol g^{-1}) > RH ($0.60\text{--}2.04$ mmol g^{-1}) > PK ($0.80\text{--}1.82$ mmol g^{-1}). Consistent with the results of the FTIR, the content of surface acid functional groups decreased with increasing pyrolysis temperature, indicating the transformation of aliphatic moieties into aromatic moieties (Chun et al., 2004). The results suggested that biochar types charred at 300 $^{\circ}\text{C}$ have larger amount of acid functional groups compared to when charred at 650 $^{\circ}\text{C}$.

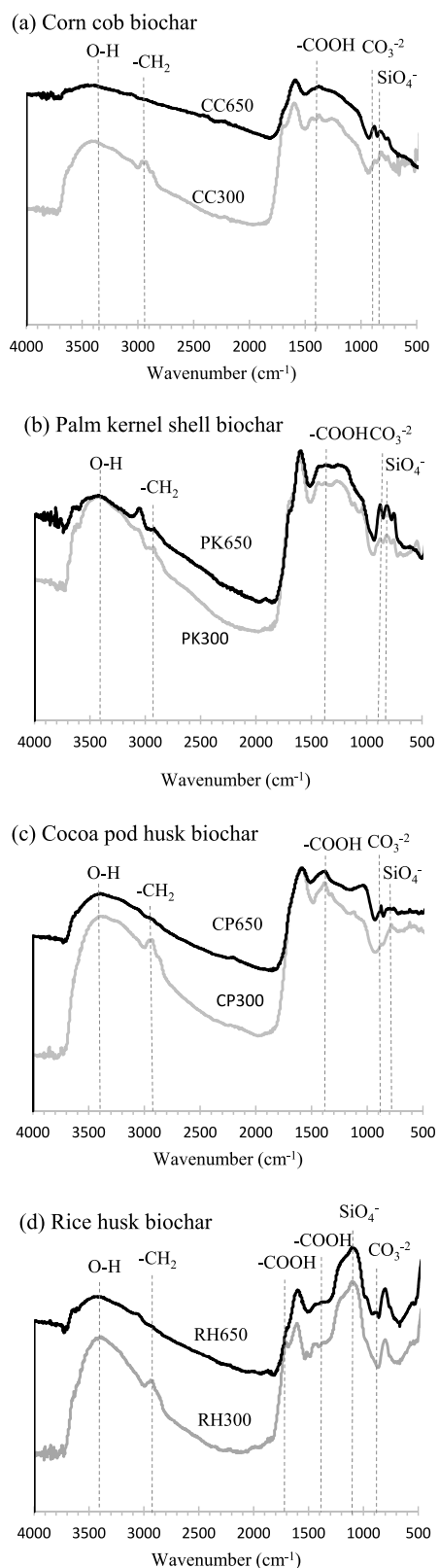


Fig. 1. Photoacoustic spectroscopy-FTIR analysis of (a) corn cob biochar (CC) charred at 300 °C (CC300) and 650 °C (CC650), (b) palm kernel shell biochar (PK) charred at 300 °C (PK300) and 650 °C (PK650), (c) cocoa pod husk biochar (CP) charred at 300 °C (CP300) and 650 °C (CP650), rice husk biochar (RH) charred at 300 °C (RH300) and 650 °C (RH650).

Table 2

Kinetic parameters related to the adsorption of phosphate onto corn cob biochar (CC300 and CC650), palm kernel shell biochar (PK300 and PK650), cocoa pod husk biochar (CP300 and CP650) and rice husk biochar (RH300 and RH650).

Kinetics models	Biochar							
	CC300	CC650	PK300	PK650	CP300	CP650	RH300	RH650
Experimental Q_e (mg g ⁻¹)	3.6	6.5	5.0	9.4	3.0	6.0	3.9	7.6
Pseudo-first-order								
Calculated q_e (mg g ⁻¹)	2.8	4.5	3.5	10.9	2.6	5.4	2.6	4.7
K_1 (h ⁻¹)	0.21	0.18	0.17	0.38	0.47	0.37	0.26	0.46
R^2	0.94	0.76	0.94	0.90	0.85	0.79	0.87	0.85
Pseudo-second-order								
Calculated q_e (mg g ⁻¹)	4.2	7.7	5.2	10.3	2.9	6.2	4.2	8.2
K_2 (mg g ⁻¹ h ⁻¹)	0.11	0.05	0.12	0.06	0.22	0.09	0.15	0.08
R^2	0.99	0.98	0.99	0.99	0.99	0.99	0.99	0.99
Elovich								
α (mg g ⁻¹ h ⁻¹)	4.14	3.25	8.85	22.94	39.25	14.27	5.51	8.41
β (g mg ⁻¹)	1.17	0.55	1.03	0.54	2.70	0.96	1.16	0.50
R^2	0.95	0.93	0.99	0.99	0.95	0.95	0.97	0.91
Intraparticle diffusion								
K_f (mg g ⁻¹ h ^{0.5})	0.28	0.63	0.33	0.65	0.11	0.38	0.30	0.58
C (mg g ⁻¹)	2.00	3.90	2.50	5.30	1.91	3.05	1.68	3.00
R^2	0.72	0.65	0.81	0.82	0.65	0.85	0.84	0.66

3.2. Effect of contact time on PO₄-P adsorption

Fig. 2 shows contact time effect of PO₄-P adsorption onto the biochars at the two pyrolysis temperatures, 300 °C and 650 °C. The results showed that the adsorption of PO₄-P from solution onto the biochars significantly ($p < 0.05$) increased in the first few hours and progressively remained stable until approaching an apparent equilibrium varying from 6 to 15 h. The initial rapid phase and the proceeding slow phase of PO₄-P adsorption is a commonly observed phenomenon (Saeed et al., 2005; Pelleria et al., 2012). For a porous adsorbent such as biochar, the initial rapid adsorption phase likely reflected surface adsorption while the subsequent slow phase reflected diffusion of PO₄-P into the pores. The equilibrium time for maximum PO₄-P adsorption was faster for the 300 °C-biochar types than the biochar types pyrolyzed at 650 °C. This can be explained by their higher CEC and lower Fe and Al contents (Table 1).

To predict and understand the PO₄-P kinetics adsorption mechanism, three popular chemical kinetics mathematical models were used, namely pseudo-first-order (PFO), pseudo-second-order (PSO) and Elovich models (Ho, 2006).

$$\frac{dq_t}{dt} = K_1 (q_e - q_t) \quad \text{Pseudo-first-order}$$

$$\frac{dq_t}{dt} = K_2 (q_e - q_t)^2 \quad \text{Pseudo-second-order}$$

$$\frac{dq_t}{dt} = \alpha \exp(-\beta q_t) \quad \text{Elovich}$$

Here q_t (mg g⁻¹) and q_e (mg g⁻¹) are the amount of P adsorbed at time t and at equilibrium, respectively, K_1 (h⁻¹) and K_2 (mg g⁻¹h⁻¹) are the PFO and PSO adsorption rate constants, respectively, α (mg g⁻¹h⁻¹) is the initial adsorption rate and β (g mg⁻¹) is the desorption constant.

PFO and PSO describe the kinetics of the solid-solution system based on physisorption and chemisorption, respectively, with regards to the adsorbent capacity (Ho, 2006). The Elovich model, which was formally used to elucidate the gas adsorption phenomenon, is now widely used to describe the aqueous contaminants sorption by assuming a strong heterogeneity of the sorbent surface (Plazinski et al., 2009). It is an empirical equation which accounts for the contribution of desorption (Yao et al., 2011).

PSO provided a better description of the kinetics of PO₄-P adsorption than PFO (Fig. 3, Table 2). Aside the high linear dependency of q_t on t ($R^2 = 0.99$), the calculated q_e values of the PSO were closer to the experimental values (Q_e). Thus, the adsorption of PO₄-P onto the biochars was better characterized by the PSO and thereby occurred through chemisorption (Ho, 2006). The adsorption rate constant, K_2 (mg g⁻¹h⁻¹) ranged from 0.05 to 0.22 mg g⁻¹h⁻¹, relatively lower for biochar produced at 650 °C. The higher K_2 values for the 300 °C-biochars demonstrate an increase rate of PO₄-P adsorption on these adsorbents than the 650 °C-biochars with time.

The Elovich model also fitted the data well for all the biochar types ($R^2 > 0.90$) (Table 2). The adsorption rate, α , varied among all the adsorbents (3.25 to 39.25 mg g⁻¹h⁻¹). Desorption constant, β increased in the order of RH650 < PK650 < CC650 < CP650 < PK300 < RH300 < CC300 < CP300. Obviously, the β values were relatively higher for the 300 °C-biochars than the 650 °C-biochars implying a higher desorption of adsorbed PO₄-P. With regards to the assumption underlying the Elovich model, the active surface sites on the biochars used in the study are heterogeneous and therefore exhibit different activation energy for chemisorption.

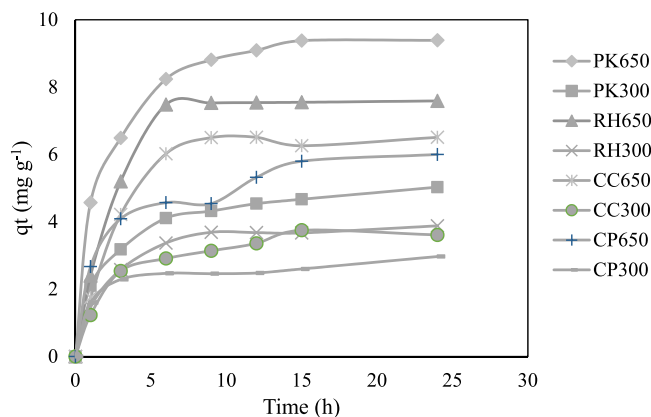


Fig. 2. Effect of contact time on phosphate onto corn cob biochar (CC300 and CC650), palm kernel shell biochar (PK300 and PK650), cocoa pod husk biochar (CP300 and CP650) and rice husk biochar (RH300 and RH650).

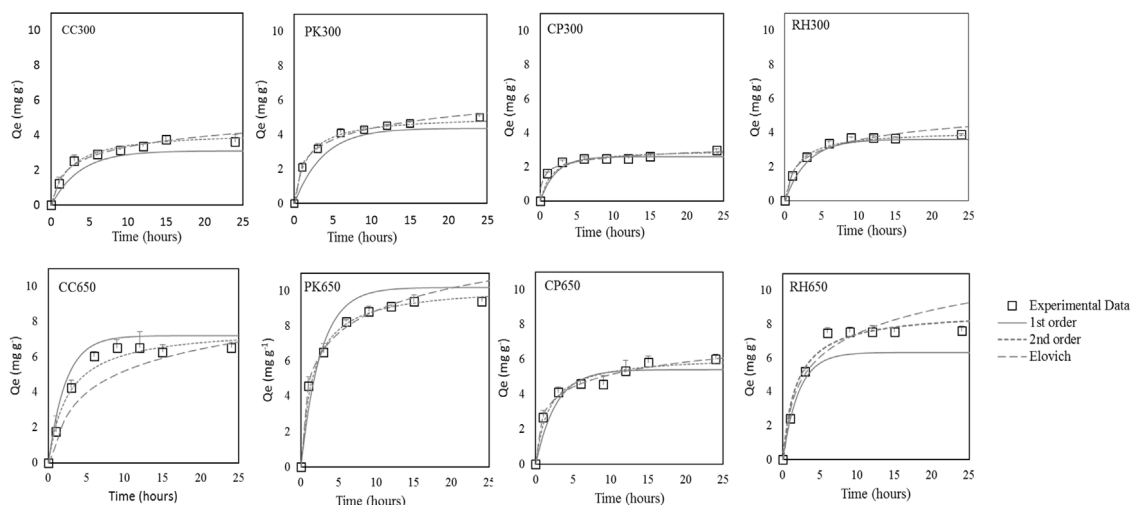


Fig. 3. Kinetics models (Pseudo-first-order, Pseudo-second-order & Elovich) of phosphate onto corn cob biochar, palm kernel shell biochar, cocoa pod husk biochar and rice husk biochar at 300 °C and 650 °C.

For a porous adsorbent such as biochar, the diffusion mechanism of $\text{PO}_4\text{-P}$ into it is of importance in the adsorption processes. Nonetheless, the chemical kinetics models cannot describe the diffusion mechanism. Intraparticle diffusion model was therefore deployed (Weber and Morris, 1963).

$$q_t = K_i t^{1/2} + C \quad \text{Intraparticle diffusion}$$

Here, K_i is the intraparticle diffusion rate constant ($\text{g } \mu\text{g}^{-1}$) and C ($\text{g } \mu\text{g}^{-1}$) is a constant associated to the boundary layer thickness, both can be obtained from the plot of q_t versus $t^{0.5}$ (Fig. 4).

$\text{PO}_4\text{-P}$ adsorbed showed a high linear dependency on the square root of time ($R^2 > 0.60$), implying that intraparticle diffusion played a role in controlling $\text{PO}_4\text{-P}$ adsorption on the biochar types. However, since the C values are greater than zero ($C > 0$) (Table 2), intraparticle diffusion is not the sole rate limiting step. The earlier and later portions of the plot are the film diffusion and intraparticle adsorption stages, respectively. The K_i values for the 650 °C-biochars was higher than the 300 °C-biochar types, which could be due to the highly porous nature of biochar produced at high temperatures (Zhang et al., 2015b). The lower values of C observed for 300 °C-biochars implies that with reference to the rate-limiting step in the overall sorption process in these biochar types, surface film diffusion was less involved.

The validity of the intraparticle model coupled with Elovich and PSO in the present study suggests that $\text{PO}_4\text{-P}$ adsorption was initially controlled by chemisorption and subsequently by intraparticle diffusion. These findings are in agreement with the results on P adsorption onto mallee biochar (Zhang et al., 2016).

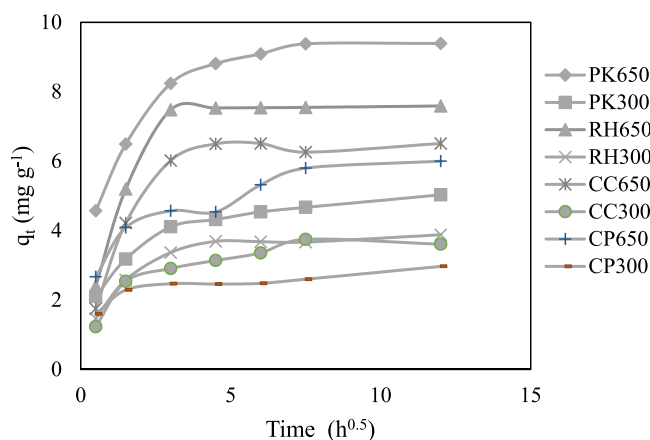


Fig. 4. Intraparticle diffusion modeling for kinetics of phosphate onto corn cob biochar (CC300 and CC650), palm kernel shell biochar (PK300 and PK650), cocoa pod husk biochar (CP300 and CP650) and rice husk biochar (RH300 and RH650).

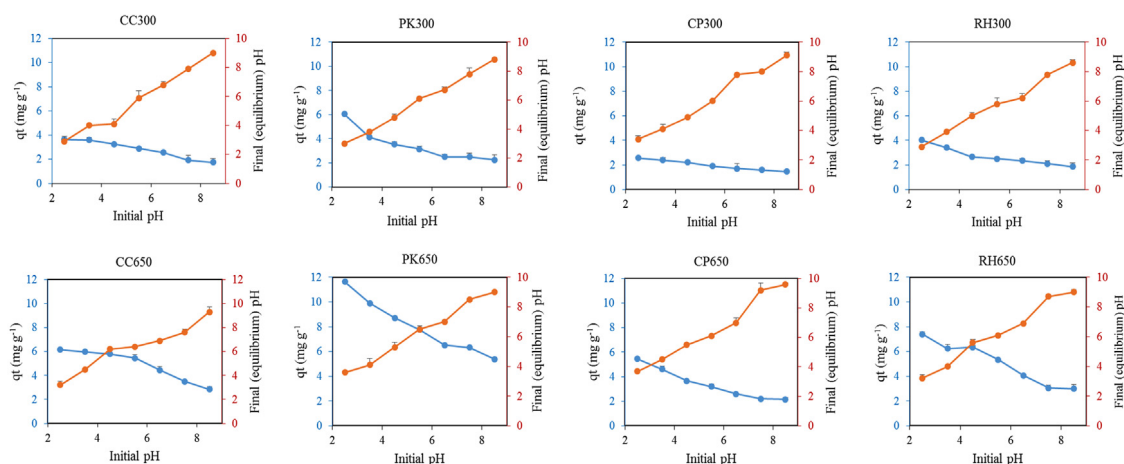


Fig. 5. Effect of pH on the adsorption of phosphate onto corn cob biochar (CC300 and CC650), palm kernel shell biochar (PK300 and PK650), cocoa pod husk biochar (CP300 and CP650) and rice husk biochar (RH300 and RH650). The red and blue colored lines show final pH (equilibrium pH) and amount of phosphate adsorbed (qt), respectively. (For interpretation of the references to color in this figure legend, the reader is referred to the web version of this article.)

3.3. Effect of pH on PO_4 -P adsorption

The amount of the PO_4 -P adsorbed onto the biochars was highly related to pH (Fig. 5). This implies that pH played a major role in the entire adsorption process, particularly on the speciation of PO_4 -P and the surface charge of the biochar and therefore its adsorption capacity. The amount of PO_4 -P adsorbed was inversely related to the equilibrium pH (2.8 to 9.6) for the biochar types (Fig. 5). Biochar mostly has low point of zero charge (PZC) (Mukherjee et al., 2011) and therefore at pH > 4.5 negative surface charges will be created resulting in low PO_4 -P sorption (Xie et al., 2014). The PAS-FTIR analysis indicated ample amounts of oxygen-containing functional groups among which is carboxylate ($-COOH$) that contributes considerably to the negative surface charge of the biochar types (Fig. 1). The surface negative potential of biochar decreases with increasing equilibrium pH (Li et al., 2017). Thus, the low affinity of PO_4 -P by the biochars at high solution pH could be a result of the increased electrostatic anion repulsion between the negatively charged biochar surfaces and the negatively charged PO_4 -P ions, and also because of increased competitive effect of OH^- ion.

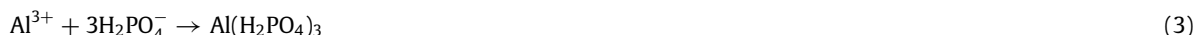
The equilibrium pH for maximum PO_4 -P adsorption onto the biochar types was between 2.6 and 4.8. This was in line with equilibrium pH i.e. 3.0 to 6.5 reported for maximum PO_4 -P sorption by potato, canola and walnut shell biochar (Feizi and Jalali, 2014). At the observed equilibrium pH (2.6 to 4.8) of the present study, Fe and Al would be active in solution. However, Ca and Mg would be precipitating or polymerizing out of solution. The presence of Fe and Al, may in part explain PO_4 -P adsorption on the biochar types. This collaborated with the results of Shepherd et al. (2017), mapping the

mechanism of P capture in biochar and also [Sibrell et al. \(2009\)](#), examining $\text{PO}_4\text{-P}$ sorption in ochres biochar with differing chemical compositions, that $\text{PO}_4\text{-P}$ removal from solution was controlled by Fe and Al. Similarly, $\text{PO}_4\text{-P}$ adsorption on Al and Fe (hydr)oxide-biochar has been reported ([Manning and Goldberg, 1996](#)). The PK650 and RH650 biochar types with relatively high Al and Fe concentration had high sorption capacities. The acid functional groups (e.g. $-\text{COOH}$ and $-\text{OH}$) are retained at 300°C but tended to decrease at 650°C ([Table 1](#) & [Fig. 1](#)), culminating in less negative charges and increased PZC at high temperatures ([Li et al., 2017](#)). As a consequence, low pyrolysis biochars had a higher CEC relative to high pyrolysis biochars. With higher CEC, it is a matter of course that these biochar types would have the highest repulsion for the $\text{PO}_4\text{-P}$ anions which would translate into the least adsorption. The variation in the sorption capacity of the biochar types under the studied pH range suggests differences in surface functional properties of these biochar types as influenced by pyrolysis temperature and feedstock.

3.4. Changes in pH and mechanism of $\text{PO}_4\text{-P}$ adsorption

To fully comprehend the mechanisms of $\text{PO}_4\text{-P}$ adsorption onto the biochar types, the ΔpH values (difference between equilibrium pH of biochar sample with $\text{PO}_4\text{-P}$ solution and equilibrium pH of blank) were calculated ([Table 3](#)). In accordance with the ΔpH values, there was no change in pH with decreasing $\text{PO}_4\text{-P}$ adsorption for PK650, RH300, RH650, CP300, CC300 and CC650 since the ΔpH were below 0.5 pH units ([Nartey et al., 2000](#)). However, CP650 and PK300 showed respective increase and decrease in pH which were more than pH 0.5 units. The 0.5 units increase in the ΔpH of CP650 coincided with the maximum $\text{PO}_4\text{-P}$ sorption capacity ([Table 3](#)), indicating the release of hydroxyl ions (OH^-) during the adsorption process. Similarly, the maximum $\text{PO}_4\text{-P}$ adsorption of PK300 coincided with a 0.7 unit decrease in pH ([Table 3](#)), implying the release of protons (H^+ or H_3O^+) during the adsorption process.

From the kinetics adsorption study, $\text{PO}_4\text{-P}$ adsorption onto the biochars occurred predominantly through chemisorption ([Table 2](#)). The maximum $\text{PO}_4\text{-P}$ sorption for CP300, RH650 and CC650 occurred at equilibrium pH from 3.2 to 3.4. Within this equilibrium pH range (3.2 to 3.4), Fe and Al from the ash in the biochar will exist as free Fe^{3+} and Al^{3+} ions ([Marion et al., 1976](#)). At this equilibrium pH range, 95% of $\text{PO}_4\text{-P}$ exists as H_2PO_4^- and only 5% as H_3PO_4 since pK_1 of $\text{PO}_4\text{-P}$ is 2.15 ([Lindsay, 1979](#)). Surface precipitation of adsorbed $\text{PO}_4\text{-P}$ with Fe and Al on CP300, RH650 and CC650 will thus occur (Eqs. (2) and (3)).



Phenol group ($-\text{OH}$) was identified on CP300, RH650 and CC650 and can remove $\text{PO}_4\text{-P}$ through ligand exchange. Maximum adsorption for CP300, RH650 and CC650 were observed between equilibrium pH 3.2 and 3.4. These pHs are 5.6 to 5.8 pH units below the pK_a of $-\text{OH}$ ($\text{pK}_a = 9.0$) ([Evangelou, 1998](#)). Thus, the $-\text{OH}$ functional groups would be protonated. At these equilibrium pHs, the dominant $\text{PO}_4\text{-P}$ species in solution is H_2PO_4^- . $\text{PO}_4\text{-P}$ removal from the solution by the protonated phenol may occur via ligand exchange (Eq. (4)) and electrostatic attraction (Eq. (5)):



Because there were no changes in solution pH after adsorption, Eqs. (4) and (5) are also plausible. In Eq. (3), the $\text{PO}_4\text{-P}$ is retained and could be exchanged by other anions (e.g. Cl^- , NO_3^-) back into solution. In Eq. (4), however, structural $-\text{OH}$ is replaced and the $\text{PO}_4\text{-P}$ is locked up. This mechanism may be ideal for water cleansing.

Carboxylic ($-\text{COOH}$), a functional group with pK_a value of 4.8 (Sparks 1995) was found on RH300 and CC300. The equilibrium pHs for maximum $\text{PO}_4\text{-P}$ adsorption were 2.6 and 2.7 for RH300 and CC300, respectively. The equilibrium pHs (2.6 and 2.7) were more than 2 pH units below the pK_a of $-\text{COOH}$ and therefore it will be protonated. The removal of $\text{PO}_4\text{-P}$ by the protonated $-\text{COOH}$ can occur via electrostatic attraction (Eq. (6)). Again, with no change in equilibrium pH as a result of adsorption, Eq. (6) is plausible. $\text{PO}_4\text{-P}$ adsorbed by RH300 and CC300 is weakly held (electrostatic attraction) and can, therefore, easily be exchanged by other anions in the solution.



The maximum $\text{PO}_4\text{-P}$ adsorption capacity of PK650 occurred at equilibrium pH of 3.6. At this pH, and with the presence of Fe^{3+} and Al^{3+} , there could be species of FeOH_2^+ and AlOH_2^+ in the solution ([Marion et al., 1976](#)). There was also no change in equilibrium pH after adsorption. Thus, in addition to surface precipitation (Eqs. (2) and (3)), PK650 can also adsorb $\text{PO}_4\text{-P}$ by ligand exchange (Eqs. (7) and (8)).



The equilibrium pH (4.1) at which maximum $\text{PO}_4\text{-P}$ adsorption occurred on PK300 coincided with a 0.7 unit decrease in equilibrium pH, implying a release of proton after $\text{PO}_4\text{-P}$ adsorption. At the equilibrium pH of 4.1, Fe^{3+} , Al^{3+} , FeOH_2^+ , AlOH_2^+ , FeOH^{2+} and AlOH^{2+} all exist in solution monomeric forms of Fe and Al ([Marion et al., 1976](#)). At this equilibrium

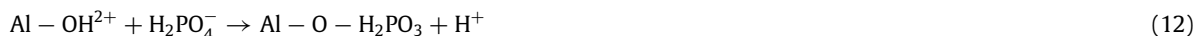
Table 3

Changes in pH and amount of phosphate adsorbed by the biochar types.

Corn cob biochar				Palm kernel shell biochar				Cocoa pod husk biochar				Rice husk biochar			
CC300		CC650		PK300		PK650		CP300		CP650		RH300		RH650	
ΔpH^a	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})	ΔpH	$\text{PO}_4\text{-P}$ adsorbed (mg g^{-1})
−0.2	3.64	0.0	5.99	0.1	4.11	0.0	11.63	0.3	2.57	0.3	4.60	−0.3	4.02	−0.2	7.40
0.3	3.59	0.2	6.25	−0.7	6.05	−0.1	9.90	0.2	2.39	0.5	5.45	−0.2	3.40	−0.2	6.25
−0.2	3.25	−0.2	5.79	−0.1	3.54	−0.1	8.72	−0.1	2.22	0.1	3.65	0.2	2.65	0.2	6.35
−0.3	2.90	−0.3	5.46	−0.2	3.15	−0.1	7.77	−0.3	1.89	−0.1	3.20	−0.1	2.50	−0.4	5.35
−0.1	2.55	−0.1	4.45	−0.2	2.51	−0.1	6.53	0.4	1.70	0.0	2.60	0.3	2.35	0.2	4.06
0.1	1.92	−0.3	3.50	0.1	2.50	0.1	6.32	−0.3	1.60	−0.3	2.20	−0.1	2.10	0.3	3.05
0.3	1.75	0.4	2.85	0.0	2.25	−0.2	5.38	−0.1	1.45	0.2	2.15	0.4	1.85	0.2	3.00

^aThe difference between equilibrium pH of biochar sample with $\text{PO}_4\text{-P}$ solution and equilibrium pH of blank without $\text{PO}_4\text{-P}$ solution.

pH (4.1), 95% of $\text{PO}_4\text{-P}$ exist as H_2PO_4^- and 5% as H_3PO_4 (Lindsay, 1979). $\text{PO}_4\text{-P}$ adsorption by PK300 can, therefore, occur by precipitation and ligand exchange with a concomitant release of protons (Eqs. (9)–(14)).



However, as the decrease in pH was only 0.7 pH units, it is unlikely that Eqs. (9) and (10) occurred since these equations depict the release of 2 moles of H^+ that would have accounted for a greater decrease in pH. The more plausible reactions are thus Eqs. (11)–(14).

The maximum $\text{PO}_4\text{-P}$ adsorption onto CP650 occurred at an equilibrium pH of 3.2. At this pH the $-\text{COOH}$ will exist mainly in its neutral form. $\text{PO}_4\text{-P}$ adsorption by $-\text{COOH}$ can occur via ligand exchange with a concomitant release of OH^- ions (Eq. (15)).



In summary, the mechanisms of $\text{PO}_4\text{-P}$ adsorption onto the biochar types were diversiform, mainly including electrostatic attraction, surface precipitation and ligand exchange; which depended largely on the feedstock and the pyrolysis temperature.

4. Conclusion

The study indicated that the biochar types have affinity for $\text{PO}_4\text{-P}$ adsorption. $\text{PO}_4\text{-P}$ adsorption was faster on the low pyrolysis biochar types (300 °C) as compared to their counterparts (650 °C). Increasing pyrolysis temperature proportionally increased the $\text{PO}_4\text{-P}$ adsorption. Maximum $\text{PO}_4\text{-P}$ adsorption on the biochar types occurred at low pHs (2.6–4.8). The adsorption of $\text{PO}_4\text{-P}$ onto the heterogeneous surface of the biochar types was predominantly controlled by intraparticle diffusion and chemisorption (electrostatic attraction, surface precipitation and ligand exchange). The use of RH300 and CC300 present an interesting option for waste recycling as a slow $\text{PO}_4\text{-P}$ releasing fertilizer. PK, CP, CC650 and RH650 can also serve as a strong $\text{PO}_4\text{-P}$ adsorbent in tertiary wastewater treatment since the $\text{PO}_4\text{-P}$ adsorbed (surface precipitation and ligand exchange) cannot easily be desorbed.

Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

CRediT authorship contribution statement

Joseph Osafo Eduah: Conceptualization, Methodology, Investigation, Visualization, Formal analysis, Writing - original draft. **Eric Kwesi Nartey:** Supervision, Methodology, Writing - review & editing. **Mark Kofi Abekoe:** Supervision, Methodology, Writing - review & editing. **Stephan Weck Henriksen:** Investigation, Resources, Formal analysis. **Mathias Neumann Andersen:** Writing - review & editing, Funding acquisition.

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References

- Almeelbi, T., Bezbaruah, A., 2012. Aqueous phosphate removal using nanoscale zero-valent iron. *J. Nanopart. Res.* 14, 1–14.
- Biswas, B.K., Inoue, K., Ghimire, K.N., Harada, H., Ohto, K., Kawakita, H., 2008. Removal and recovery of phosphorus from water by means of adsorption onto orange waste gel loaded with zirconium. *Bioresour. Technol.* 99, 8685–8690.
- Boehm, H.P., 1994. Some aspects of the surface chemistry of carbon blacks and other carbons. *Carbon* 32, 759–769.
- Bourke, J., Manley-Harris, M., Fushimi, C., Dowaki, K., Nunoura, T., Antal, M.J., 2007. Do all carbonized charcoals have the same chemical structure? A model of the chemical structure of carbonized charcoal. *Ind. Eng. Chem. Res.* 46, 5954–5967.

- Brewer, C.E., Schmidt-Rohr, K., Satrio, J.A., Brown, R.C., 2009. Characterization of biochar from fast pyrolysis and gasification systems. *Environ. Progress. Sustain. Energy* 28, 386–396.
- Chun, Y., Sheng, G.Y., Chiou, C.T., Xing, B.S., 2004. Compositions and sorptive properties of crop residue-derived chars. *Environ. Sci. Technol.* 38, 4649–4655.
- Demirbas, A., 2008. Production and characterization of bio-chars from biomass via pyrolysis. *Energy Sources A* 28, 413–422.
- Dodds, W.K., Bouska, W.W., Eitzmann, J.L., Pilger, T.J., Pitts, K.L., Riley, A.J., Schloesser, J.T., Thornbrugh, D.J., 2008. Eutrophication of US freshwaters: Analysis of potential economic damages. *Environ. Sci. Technol.* 43, 12–19.
- Evangelou, V.P., 1998. *Environmental Soil and Water Chemistry: Principles and Applications*. John Wiley and Sons, Inc., Canada.
- Feizi, M., Jalali, M., 2014. Sorption of aquatic phosphorus onto native and chemically-modified plant residues: modeling the isotherm and kinetics of sorption process. *Desalin. Water Treat.* 45, 121–133.
- Ho, Y.S., 2006. Review of second-order models for adsorption systems. *J. Hazard. Mater.* 136, 681–689.
- Jiang, J., Yuan, M., Renkou, X., Bish, D.L., 2015. Mobilization of phosphate in variable-charge soils amended with biochars derived from crop straws. *Soil Tillage Res.* 146, 139–147.
- John, M.K., 1970. Colorimetric determination of phosphorus in soil and plant materials with ascorbic acid. *Soil Sci.* 109 (4), 214–220.
- Jung, K.W., Ok, Y.S., Hwang, M.J., Ahn, K.H., 2015. Kinetic study on phosphate removal from aqueous solution by biochar derived from peanut shell as renewable adsorptive media. *Int. J. Environ. Sci. Technol.* 12, 3363–3372.
- Kasozzi, G.N., Zimmerman, A.R., Nkedi-Kizza, P., Gao, B., 2010. Catechol and humic acid sorption onto a range of laboratory-produced black carbons (Biochars). *Environ. Sci. Technol.* 44, 6189–6195.
- Lehmann, J., 2007. Bio-energy in the Black. *Front. Ecol. Environ.* 5, 381–387.
- Li, H., Dong, X., da Silva, E.B., de Oliveira, L.M., Chen, Y., Ma, L.Q., 2017. Mechanisms of metal sorption by biochars: Biochar characteristics and modifications. *Chemosphere* 178, 466–478.
- Lindsay, W.L., 1979. *Chemical Equilibria in Soils*. John Wiley & Sons, New York.
- Manning, B.A., Goldberg, S., 1996. Modeling competitive adsorption of arsenate with phosphate and molybdate on oxide minerals. *Soil Sci. Am. J.* 60, 121–131.
- Marion, G.M., Hendricks, D.M., Dutt, G.R., Filler, W.H., 1976. Aluminum and silica solution in soils. *Soil Sci.* 127, 76–85.
- Mukherjee, A., Zimmerman, A.R., Harris, W., 2011. Surface chemistry variations among a series of laboratory-produced biochars. *Geoderma* 163, 247–255.
- Nartey, E., Matsue, N., Henmi, T., 2000. Adsorptive mechanism of orthosilicic acid on nano ball allophane. *Clay Sci.* 11, 125–136.
- Pellera, F.M., Giannis, A., Kalderis, D., Anastasiadou, K., Stegmann, R., Wang, J.Y., Gidarakos, E., 2012. Adsorption of Cu (II) ions from aqueous solutions on biochars prepared from agricultural byproducts. *J. Environ. Manag.* 96, 35–42.
- Plazinski, W., Rudzinski, W., Palzinska, A., 2009. Theoretical models of sorption kinetics including a reaction mechanism: a review. *Adv. Colloid Interface Sci.* 152, 2–13.
- Saeed, A., Akhter, M.W., Iqbal, M., 2005. Removal and recovery of heavy metals from aqueous solution using papaya wood as a new biosorbent. *Sep. Purif. Technol.* 45, 25–31.
- Shepherd, J.G., Joseph, S., Sohi, S.P., Heal, K.V., 2017. Biochar and enhanced phosphate capture: Mapping mechanisms to functional properties. *Chemosphere* 179, 57–74.
- Sibrell, P.L., Montgomery, G.A., Ritenour, K.L., Tucker, T.W., 2009. Removal of phosphorus from agricultural wastewaters using adsorption media prepared from acid mine drainage sludge. *Water Res.* 43, 2240–2250.
1995. USEPA Ecological Restoration: A Tool to Manage Stream Quality. Report EPA 841-F-95-007, US EPA, Washington, DC, USA.
- Usman, A.R.A., Abduljabbar, A., Vithanage, M., Ok, Y.S., Ahmad, M., Ahmad, M., Elfaki, J., Abdulazeem, S., Al-Wabel, M.I., 2015. Biochar production from date palm waste: Charring temperature induced changes in composition and surface chemistry. *J. Anal. Appl. Pyrolysis* 115, 392–400.
- Wahab, M.A., Hassine, R.B., Jellali, S., 2011. Removal of phosphorus from aqueous solution by *Posidonia oceanica* fibers using continuous stirring tank reactor. *J. Hard Mater.* 191, 333–341.
- Weber, W.J., Morris, J.C., 1963. Kinetics of adsorption on carbon from solution. *J. Sanit. Eng.: Div. Am. Soc. Civ. Eng.* 89, 31–60.
- Xie, M., Chen, W., Xu, Z., Zheng, S., Zhu, D., 2014. Adsorption of sulfonamides to demineralized pine wood biochars prepared under different thermochemical conditions. *Environ. Pollut.* 186, 187–194.
- Yao, Y., Gao, B., Chen, J., Yang, L., 2012. Engineered biochar reclaiming phosphate from aqueous solutions: mechanisms and potential application as a slow-release fertilizer. *Environ. Sci. Technol.* 47 (15), 8700–8708.
- Yao, Y., Gao, B., Inyang, M., Zimmerman, A.R., Cao, X.D., Pullammanappallil, P., Yang, L.Y., 2011. Removal of phosphate from aqueous solution by biochar derived from anaerobically digested sugar beet tailings. *J. Hard Mater.* 190 (1–3), 501–507.
- Zhang, H., Chen, C., Gray, E.M., Boyd, S.E., Yang, H., Zhang, D., 2016. Roles of biochar in improving phosphorus availability in soils: a phosphate adsorbent and a source of available phosphorus. *Geoderma* 276, 1–6.
- Zhang, J., Liu, J., Liu, R.L., 2015b. Effects of pyrolysis temperature and heating time on biochar obtained from the pyrolysis of straw and lignosulfonate. *Bioresour. Technol.* 176, 288–291.
- Zhang, L., Loaiciga, H.A., Xu, M., Du, C., Du, Y., 2015. Kinetics and mechanisms of phosphorus adsorption in soils from diverse ecological zones in the source area of a drinking-Water Reservoir. *Int. J. Environ. Res. Public Health* 12, 14312–14326.
- Zhao, L., Cao, X., Masek, O., Zimmerman, A., 2013. Heterogeneity of biochar properties as a function of feedstock sources and production temperatures. *J. Hard Mater.* 19, 256–257.